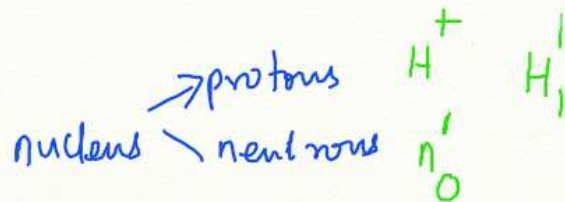
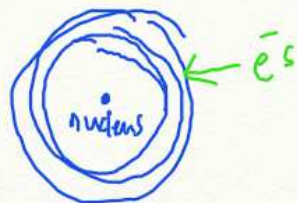


Chem. Tut. Julia

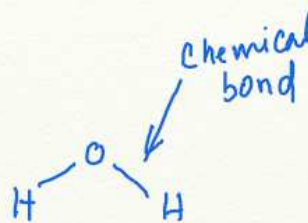
Elements - 103 of them

Oxygen : O



Compound - more than 1 element

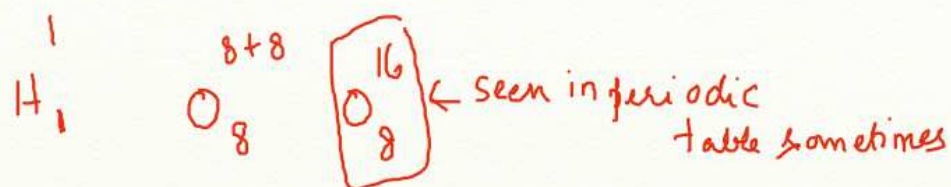
ex: H_2O



Bond: is formed due to electrons

Atomic No: no. of protons

Mass No: no of protons + neutrons



no. of protons = no. of electrons in an element

no of neutrons = mass no. - atomic no.

Ex: for O:

$$= 16 - 8 = 8 \text{ neutrons}$$

Isotopes

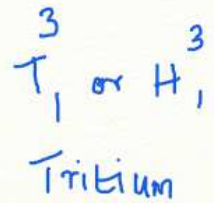
elements having same atomic no.
but different mass no.

H has 3 isotopes



or H_1^2

Deuterium



or H_1^3

Tritium

if unstable, isotopes disintegrates